Q1.Select the correct answer from the words in bracket.

1. He arranged elements in increasing order of atomic numbers. [Dobereiner/ Moseley / Mendeleeff
2. Is a metal in period 2 having electronic configuration 2, 1. [Beryllium / Lithium / Sodium]
3. Is a period having elements from atomic no. 11 to 18. [period – 1/2/3]
4. The most reactive halogen from group 17. [chlorine/fluorine/bromine]
5. Is the group number of the element whose atomic number is 4. [group - 1/group – 2/group – 18]

Q2.Fill in the blanks from the words A to F given below.

A: Decreases B: Increases C: Remains same D: Increases by one E: Electropositive F: Electronegative

Across a period from left to right in the Modern Periodic Table.

No. of electron shells ..... ; No. of valence electrons .....; Electronegativity .....;
Character of elements changes from ..... to .....

Down a group in the Modern Periodic Table.

No. of electron shells .....; No. of valence electrons .....; Electronegativity .....;
Character of elements changes from ..... to .....

Q3.Match the elements of List-I with their type from List-II.

List I	List II
1. ighly reactive, univalent metals	**A: Halogens**
2. Highly reactive, univalent non-metals	**B: Transition elements**
3. nreactive, inert elements	**C: Alkali metals**
4. lements of group 3 to 12 [IB to VIIB and VIII]	**D: Lanthanide series**
5. adioactive elements	**E: Noble gases**
	F: Actinide series

Q4.Complete electtonic configuration of following

1. Hydrogen	1	.....
2. Carbon	6	....., ...
3. Nitrogen	7	....., ...
4. Oxygen	8	....., ...
5. Sulphur	16	....., ..., ...
6. Chlorine	17	....., ..., ...

Q5.State the following.

1. The group to which the element with electronic configuration of 2, 8, 2 belongs.
2. The group from the groups 1[IA], 2[IIA], 16[VIA] and 17[VIIA] whose elements are most electronegative.
3. The group which contains highly electropositive metals including sodium.
4. The group whose elements are unreactive or inert.
5. The group which contains highly reactive electronegative non-metals including chlorine.

Q6.Name or state the following with reference to the elements of the first three periods of the periodic table.

(a) The noble gas having duplet arrangement of electrons.

(b) The noble gas having an electronic configuration 2, 8, 8.

(c) A metalloid in period 2 and in period 3.

(d) The number of electron shells in elements of period 1, period 2 and period 3.

(e) The valency of elements in group 1 [IA].

(f) The group whose elements have zero valency.

(g) An alkaline earth metal in period 3.

(h) The non-metallic element present in period 3 other than sulphur and chlorine.

(i) A non-metal in period 2 having electronic configuration 2, 6.

(j) An electrovalent compound formed between an alkali metal and a halogen.

(k) A covalent compound formed between an element in period 1 and a halogen.

(l) An alkali metal in period 3 which dissolves in water giving a strong alkali.

(m) A metal in period 3 having valency 3.

(n) The bridge elements of period 3 of group 1 [IA], 2 [IIA] and 13 [IIIA].

(o) The periods which contain the inner transition elements.

(p) The formula of the hydroxide of the element having electronic configuration 2, 8, 2.

(q) The valency of the element in period 3 having atomic number 17.

(r) A non-metal in period 2 which is tetravalent.