

PHYSICAL AND CHEMICAL CHANGES DPP : 01

- Which one of the step while burning a candle is not reversible?
(A) Melting of solid wax (B) Liquid wax changes into vapours
(C) Wax vapour burn into flame (D) All of these
- Which of the following changes cannot be reversed?
(A) Milk to paneer (B) Cold milk to hot milk
(C) Yarn to knitted sweater (D) Wet clothes to dry clothes
- Rusting is a -
(A) slow process (B) fast process (C) very fast process (D) none of these
- During all changes physical or chemical changes mass is
(A) gained (B) conserved (C) loss (D) none of these
- Melting of wax is a..... change while burning of candle is change.
(A) irreversible, reversible (B) reversible, irreversible
(C) physical, reversible (D) chemical, irreversible
- Bursting of cracker is a -
(A) Slow change (B) Fast change (C) Periodic change (D) none of these
- Which of the following changes include formation of new substances?
(A) Melting (B) Sublimation (C) Evaporation (D) Rusting
- A change in which one or new substances are formed is called -
(A) chemical change (B) physical change (C) corrosion (D) all of the above
- Earthquakes is an example of -
(A) periodic change (B) non-periodic change
(C) both periodic & non-periodic change (D) none of these
- Favorable conditions for a chemical reaction may be-
(i) When the reactants come in close contact.
(ii) When the reactants are heated.
(iii) When the reactants are exposed to light.
(iv) When the reactants are subjected to pressure.
(v) When a catalyst is used the same chemical reaction takes place at a faster rate
(A) i, ii and iii (B) ii, iii and iv (C) ii, iv and v (D) all five options
- What are the main characteristics of chemical change?
- Explain why:
(i) Printing is an irreversible change.
(ii) Ironing of cloth is reversible change.
(iii) Melting of wax is a reversible change.
- Differentiate between reversible and irreversible change.
- Give two examples of natural change and man-made change ?
- What are the main characteristics of physical change?

PHYSICAL AND CHEMICAL CHANGES DPP 2

1. Classify the following as a physical or a chemical change.

- (i) Drying of wet clothes
- (ii) Manufacture of salt from sea water
- (iii) Making of curd from milk
- (iv) Butter getting rancid
- (v) Growth of a tree (vi) Rusting of iron
- (vii) Boiling of water (viii) Burning of paper
- (ix) Freezing of water
- (x) Magnetisation of a piece of iron
- (xi) Burning of a piece of magnesium wire
- (xii) Dropping sodium in water.

2. By giving one example, explain what do you understand by the following terms?

- (i) Exothermic reactions
- (ii) Endothermic reactions

3. EXPLAIN

- a. The difference between physical and chemical change
- b. The conditions when there is a chemical change occurring

4. What do you understand by the term precipitate

PHYSICAL AND CHEMICAL CHANGES DPP-3

1. Fill in the blank spaces by choosing the correct words from the given list:

(List: liquid, hydrogen sulphide, exothermic, yellow, dirty green.)

- a) When quick lime is placed in water anreaction takes and the reaction mixture gets very hot.
- b) When red hot coke (solid state) , reacts with sulphur vapour the product, carbon disulphide is instate.
- c) Ferrous sulphate solution reacts with sodium hydroxide solution to form a.....precipitate of ferrous hydroxide,
- d) Lead nitrate crystals on heating strongly leave behind a solid residue, which is in colour.
- e) When copper sulphide is treated with hydrochloric acid it gives off a gaswhich has a foul smell.

2. Statements given below are incorrect. Write the correct statements.

- a) Lead nitrate crystals on strong heating decompose to form lead dioxide and oxygen gas.
- b) The reaction between magnesium and oxygen is an endothermic reaction.
- c) The decomposition of ammonium chloride into ammonia and oxygen is an exothermic reaction.
- d) Copper carbonate on strong heating leaves behind a residue. Which is green in colour.
- e) Ammonia gas and HCl gas react to form gaseous ammonium chloride.

3. Match the statements in column A, with those in column B.

	Column A		Column B
1	A chemical reaction which proceeds with the evolution of heat.	A	Chemical change
2	A chemical reaction which proceeds with the absorption of heat.	B	Exothermic reaction
3	Name of a change brought about by the decomposition of fruit juice.	C	Oxygen
4	Name of change brought about when current passes through an electric bulb	D	Endothermic reaction
5.	A gas given out when sodium nitrate crystals are heated strongly.	E	Physical change

4. Tick the most appropriate answer.

A. The reaction between magnesium and oxygen is:

- | | |
|----------------------------|--------------------------|
| a) An endothermic reaction | c) A catalyzed reaction |
| b) An exothermic reaction | d) A reversible reaction |

B. The decomposition of sodium bicarbonate on heating is:

- | | |
|----------------------------|---------------------------|
| a) An endothermic reaction | c) A catalyzed reaction |
| b) A reversible reaction | d) An exothermic reaction |

C. When lead nitrate crystals are strongly heated, the colour of one of the gas evolved is :

- | | |
|--------------------|------------------|
| a) Greenish yellow | c) Reddish brown |
| b) Pink | d) Light grey |

- D. When the ammonia gas reacts with hydrochloric acid gas, the ammonium chloride is formed which is in:
- a) Liquid state
 - b) Solid state
 - c) Gaseous state
 - d) Partly solid and partly gaseous state
- E. When silver nitrate solution is mixed with sodium chloride solution, a precipitate of silver chloride is formed, which is of:
- a) Silvery colour
 - b) Yellow colour
 - c) Grey colour
 - d) White colour